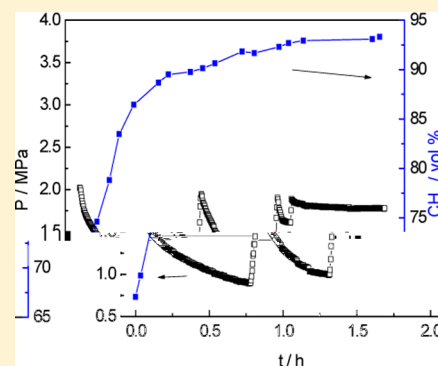


Removal of CO₂ from Biogas by Using *Tert*-Butyl Peroxy-2-ethylhexanoate and Water

Shuanshi Fan,^{*} Qi Li, Yanhong Wang, Xuemei Lang, and Jun Chen

510640, C

ABSTRACT: The effect of the addition of 2.23 g/kg of B₂O₃ to a 67.0% B₂O₃ feed on the growth and survival of rainbow trout, *Oncorhynchus mykiss*, was evaluated. The fish were fed a diet containing 67.0% B₂O₃ and 2.23 g/kg of B₂O₃ for 12 weeks. The fish fed the 67.0% B₂O₃ diet had a significantly higher survival rate (89.6%) than the fish fed the 67.0% B₂O₃ diet (26.0%). The fish fed the 67.0% B₂O₃ diet had a significantly higher growth rate (0.91 g/g) than the fish fed the 67.0% B₂O₃ diet (0.17 g/g). The fish fed the 67.0% B₂O₃ diet had a significantly higher feed conversion ratio (FCR) (1.70) than the fish fed the 67.0% B₂O₃ diet (2.23). The fish fed the 67.0% B₂O₃ diet had a significantly higher water quality index (WQI) (93.3%) than the fish fed the 67.0% B₂O₃ diet (26.0%).



1. INTRODUCTION

[illegible]

(C_4) (CO_2) 19
 CO_2 20 CO_2 21,22 C_4 23,24
 CO_2 25,26 C_4-CO_2 27
 $()$, 28 (C) , 29 (BAB) , 30 $(B \bigcirc)$, 31,32 33,34 $()$

Submitted: 12/15/2015

R : D : 5, 2015

Received: 29, 2016

A : 11, 2016

P b_i : 11, 2016

the total flow rate of the gas mixture. A flow rate of 20 L/min was maintained throughout the experiment. The gas mixture was composed of CO_2 and C_4 in a molar ratio of 0.67:0.33. The total flow rate was controlled by a mass flow controller (MFC) and a pressure transducer. The gas mixture was then passed through a series of filters and a drying column to remove any moisture. The gas mixture was then passed through a series of filters and a drying column to remove any moisture. The gas mixture was then passed through a series of filters and a drying column to remove any moisture.

2. EXPERIMENTAL SECTION

2.1. Materials.

The gas mixture was composed of CO_2 and C_4 in a molar ratio of 0.67:0.33. The total flow rate was controlled by a mass flow controller (MFC) and a pressure transducer.

Tab 1. Experimental conditions and parameters.

Parameter	Value
C_4/CO_2 molar ratio	0.67/0.33
Flow rate (L/min)	20
Pressure (atm)	1.0
Temperature (°C)	25

The gas mixture was composed of CO_2 and C_4 in a molar ratio of 0.67:0.33. The total flow rate was controlled by a mass flow controller (MFC) and a pressure transducer. The gas mixture was then passed through a series of filters and a drying column to remove any moisture. The gas mixture was then passed through a series of filters and a drying column to remove any moisture.

2.2. Apparatus and Methods.

The experimental setup consisted of a gas cylinder, a mass flow controller (MFC), a pressure transducer, and a series of filters and a drying column. The gas mixture was then passed through a series of filters and a drying column to remove any moisture. The gas mixture was then passed through a series of filters and a drying column to remove any moisture.

The gas mixture was composed of CO_2 and C_4 in a molar ratio of 0.67:0.33. The total flow rate was controlled by a mass flow controller (MFC) and a pressure transducer. The gas mixture was then passed through a series of filters and a drying column to remove any moisture. The gas mixture was then passed through a series of filters and a drying column to remove any moisture.

The gas mixture was composed of CO_2 and C_4 in a molar ratio of 0.67:0.33. The total flow rate was controlled by a mass flow controller (MFC) and a pressure transducer. The gas mixture was then passed through a series of filters and a drying column to remove any moisture. The gas mixture was then passed through a series of filters and a drying column to remove any moisture.

2.3. Calculation Methods. Number of Moles of Gas Consumed.

The number of moles of gas consumed was calculated using the following equation:

$$\Delta n = n_0 - n_t = \frac{P_0 V_0}{z_0 RT} - \frac{P_t V_t}{z_t RT} \quad (1)$$

The number of moles of gas consumed was calculated using the following equation:

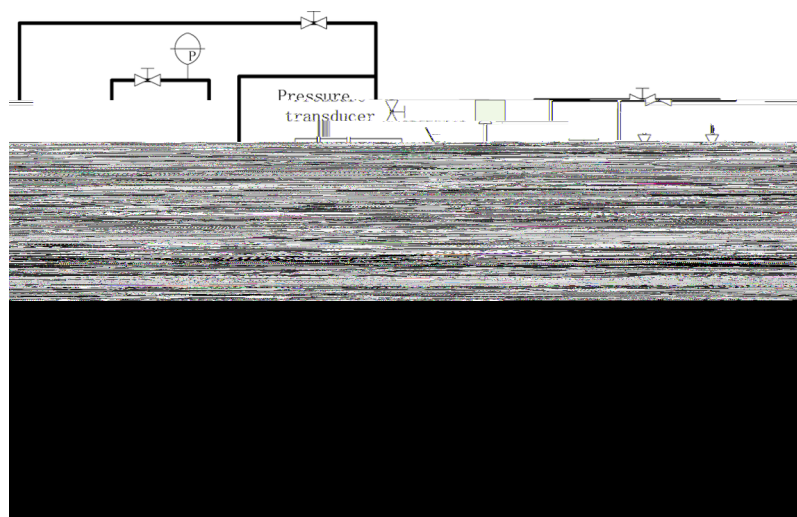


Fig. 1. Schematic diagram of the experimental setup.

The $\text{C}_{40}\text{H}_{78}\text{O}_2$ compound was identified as 2,2,4,4-tetramethyl-3,3,5,5-tetraphenyl-1,3-dioxane (BAB) by comparing its mass spectrum with the reference spectrum of BAB. The mass spectrum of the sample showed a molecular ion peak at m/z 382.1, which is consistent with the molecular weight of BAB. The fragmentation pattern of the sample was also similar to that of BAB, with major peaks at m/z 382.1, 367.1, 352.1, 337.1, 322.1, 307.1, 292.1, 277.1, 262.1, 247.1, 232.1, 217.1, 202.1, 187.1, 172.1, 157.1, 142.1, 127.1, 112.1, 97.1, 82.1, 67.1, 52.1, 37.1, 22.1, 7.1, and 0.1. The mass spectrum of the sample was therefore identified as BAB.

... 35.0% t 66.6%
 ... /
 C_4 ...
 C_{44}

...³⁶
 C_4 ...
 CO_2 ...
 C_4 ... /
 13.60.
 CO_2 ...
 C_4 ... / 5.18,
 C_4 ...
 CO_2 ... 0.5 ... /
 2.23. ... /
 C_4 ...
 C_4 ...
 3 ... CO_2 ... 26 t %
 $B + C_4 / CO_2$... /
 287 ... 2.0 ...
 ... /

- (15) ... 4A ... 13 ... 2011, 3, 314.
- (16) ... C; A ... CO₂ A ... 2013, 2, 3928.
- (17) ... C. A; B ... C ... 2014, 2, 6688.
- (18) ... C ... 2014, 2, 5745.
- (19) A ... D ... 2011, 3, 1, 2020.
- (20) ... D ... B ... 2009, 3, 20.
- (21) D ... C; O ... CO₂ C ... 2009, 2, 5603.
- (22) ... B ... A; ... CO₂-C₄ ... 2011, 01, 171.
- (23) ... D; C; B ... CO₂-C₄ ... 2013, 2, 899.
- (24) ... D; D; C; D ... 2013, 2, 7252.
- (25) ... D; C; B ... CO₂ ... 2014, 122, 206.
- (26) ... BAB ... BA ... 2009, 2, 4202.
- (27) ... A; ... D; ... 2012, 62.
- (28) ... D ... CO₂ + C₄ ... 2014, 4110.
- (29) ... CO₂ ... 2010, 1, 202.
- (30) ... CO₂/C₄ ... (B ... C ... 2013, 8, 3137.
- (31) D ... C ... CO₂ ... C₄ ... 2014, 3, 1234.
- (32) C ... D ... 2013, 8, 3494.
- (33) D ... CO₂ ... 2009, 8, 5802.
- (34) ... C ... B ... CO₂ ... 2012, 123.
- (35) ... CO₂ ... B ... 1-B t -3- ... 2015, 2, 3143.
- (36) ... B; D ... D; B ... 1997, 1090.